

## **GUIDELINES FOR LAB SUBMISSIONS**

**NAME:** Your Names

**TITLE:** Periodicity and chemical reactivity

### **INTRODUCTION**

Five paragraph intro on:

1. solubility and reactivity
2. group 2A families, who is in them what do they do
3. reactivity of members of the group 2A family
4. observations on solubility of the four compounds
5. in this study .....
  - a. Determine the relationship between...
  - b. We hypothesized that...
  - c. Our results indicate that...

### **MATERIALS METHODS**

- a) Write out your experimental design.
- b) What were your controls?
- c) What was measured?
- d) What is the independent variable for this experiment?
- e) What is the dependent variable for this experiment?
- f) Write out all mathematical equations used to make serial dilutions?
- g) How many trials were done?

### **RESULTS**

- a) Provide all of your data in this section.
- b) 1 paragraph explaining each figure (all figures at the back).
- c) You must have transition statements to move the reader through the

### **CONCLUSION:**

- a) Explain how your data either matches or contradicts your hypothesis.
- b) Explain how the results compare to placement on the periodic chart.

<b><math>Na_2SO_4</math></b>		<b>1</b>	<b>2</b>	<b>3</b>	<b>4</b>	<b>5</b>	<b>6</b>	<b>7</b>	<b>8</b>	<b>9</b>
<b><math>Mg^{+2}</math></b>	<b>A</b>									
<b><math>Ca^{+2}</math></b>	<b>B</b>									
<b><math>Sr^{+2}</math></b>	<b>C</b>									
<b><math>Ba^{+2}</math></b>	<b>D</b>									

**Figure 1:** (provide a meaning full title for this figure: explain what is it? why did you do it? what were the results?)

↑P = high level precipitant

- P = precipitant

↓P = low level precipitant

☒ = no reaction

<b><math>\text{Na}_2\text{CO}_3</math></b>		<b>1</b>	<b>2</b>	<b>3</b>	<b>4</b>	<b>5</b>	<b>6</b>	<b>7</b>	<b>8</b>	<b>9</b>
<b><math>\text{Mg}^{+2}</math></b>	<b>A</b>									
<b><math>\text{Ca}^{+2}</math></b>	<b>B</b>									
<b><math>\text{Sr}^{+2}</math></b>	<b>C</b>									
<b><math>\text{Ba}^{+2}</math></b>	<b>D</b>									

**Figure 2:** (provide a meaning full title for this figure: explain what is it? why did you do it? what were the results?)

↑P = high level precipitant

- P = precipitant

↓P = low level precipitant

☒ = no reaction

<b><math>\text{Na}_2\text{C}_2\text{O}_4</math></b>		<b>1</b>	<b>2</b>	<b>3</b>	<b>4</b>	<b>5</b>	<b>6</b>	<b>7</b>	<b>8</b>	<b>9</b>
<b><math>\text{Mg}^{+2}</math></b>	<b>A</b>									
<b><math>\text{Ca}^{+2}</math></b>	<b>B</b>									
<b><math>\text{Sr}^{+2}</math></b>	<b>C</b>									
<b><math>\text{Ba}^{+2}</math></b>	<b>D</b>									

**Figure 3:** (provide a meaning full title for this figure: explain what is it? why did you do it? what were the results?)

↑P = high level precipitant

- P = precipitant

↓P = low level precipitant

☒ = no reaction

<b>Salt</b>		<b><math>Na_2SO_4</math></b>	<b><math>Na_2CO_3</math></b>	<b><math>Na_2C_2O_4</math></b>
<b><math>Mg^{+2}</math></b>	<b>A</b>			
<b><math>Ca^{+2}</math></b>	<b>B</b>			
<b><math>Sr^{+2}</math></b>	<b>C</b>			
<b><math>Ba^{+2}</math></b>	<b>D</b>			

**Figure 4:** (provide a meaning full title for this figure: explain what is it? why did you do it? what were the results?)

- ↑R = High reactive
- R = Reactive
- ↓R = low level reactive
- ☒ = no reaction